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key , source:transatlantic-news.com You will need to comprehend how to project cash flow. Whatever your business planning objectives, cash flow is the resource in the organization, and managing cash is the one most important small business function.

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reaction products. As with many
chemistry-oriented things, the skill of
predicting reaction products is useless if
you don't actually use it for something.
Actually, that's true for just about
everything, whether chemistry-oriented

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or not. And some things are just plain ...

Steps to Predicting the Products of Chemical Reactions

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answers is correct, and the reagent that
leads to this answer is the limiting
reagent Solutions to the Molarity
Practice Worksheet For the first five
problems, you need to use the equation
that says that the Stoichiometry

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Predicting Reaction Products - Solutions

Balance the equations and predict the
products for the following reactions: 1) $3 \text{Na} + 1 \text{FeBr}_3 \rightarrow 3 \text{NaBr} + 1 \text{Fe}$ 2) 2NaOH

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+ 1 H₂SO₄ 1 Na₂SO₄ + 2 H₂O 3) 1
C₂H₄O₂ + 2 O₂ 2 CO₂ + 2 H₂O 4)
1 NH₃ + 1 H₂O 1 NH₄OH 5) 1 PbSO₄

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substance like sodium chloride and then look it up.

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place at all? With this worksheet and a solubility table, they'll understand in no time! Reaction Products Worksheet: Your students will practice making predictions about the products of chemical reactions, given only the reagents.

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Predicting Reaction Products - Solutions

Balance the equations and predict the
products for the following reactions: 1) 3

$\text{Na} + 1 \text{FeBr}_3 \rightarrow 3 \text{NaBr} + 1 \text{Fe}$ 2) 2NaOH

$+ 1 \text{H}_2\text{SO}_4 \rightarrow 1 \text{Na}_2\text{SO}_4 + 2 \text{H}_2\text{O}$ 3) 1

$\text{C}_2\text{H}_4\text{O}_2 + 2 \text{O}_2 \rightarrow 2 \text{CO}_2 + 2 \text{H}_2\text{O}$ 4)

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1 NH₃ + 1 H₂O 1 NH₄OH 5) 1 PbSO₄

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PREDICTING REACTION PRODUCTS:

COMBUSTION REACTIONS! A

hydrocarbon and oxygen (O₂) indicate a
combustion reaction ! If it is combustion,
then just write H₂O and CO₂ as

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products ! Then, balance the equation
(can be tricky) ! Examples: !
 $\underline{\quad} \text{CH}_4 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{CO}_2 + \underline{\quad} \text{H}_2\text{O}$
 $\underline{\quad} \text{C}_4\text{H}_{10} + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{CO}_2 + \underline{\quad} \text{H}_2\text{O}$

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Predicting reaction products isn't ever done. ... If the hypothetical products of your chemical compounds aren't valid (something like CaN or the like), then

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your answer will be wrong. Make sure when you're done that anything you write down is a valid and reasonable formula.

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reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be CO_2 and H_2O (assuming a complete reaction).

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